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\u201cAcids,Bases and salts\u201c How Strong are

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In the chapter on acids and bases, we saw two more definitions of acids: a compound that donates a proton (a hydrogen ion, H^+) to another compound is called a Brønsted-Lowry acid, and a Lewis acid is any species that can accept a pair of electrons. Explain why the introductory definition is a macroscopic definition, while the Brønsted-Lowry definition and the Lewis definition are microscopic definitions.

15.2 Lewis Acids and Bases - Chemistry

In the chapter on acids and bases, we saw two more definitions of acids: a compound that donates a proton (a hydrogen ion, H^+) to another compound is called a Brønsted-Lowry acid, and a Lewis acid is any species that can accept a pair of electrons. Explain why the introductory definition is a macroscopic definition, while the Brønsted-Lowry definition and the Lewis definition are microscopic definitions.

15.2 Lewis Acids and Bases - General Chemistry 1 & 2

An acid-base reaction occurs in aqueous solution between hydrochloric acid, a strong acid that completely dissociates to produce H_3O^+ , and sodium hydroxide, a strong base that completely dissociates to produce OH^- . The formula equation for this reaction is written as follows.

$$HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$$

CHAPTER 15 Acids and Bases

15.1 Acids and bases

acid derived from Latin acidus (meaning sour or tart) related to Latin acetum (meaning vinegar) characteristic properties associated with acid:

1. sour taste
2. change the color of

litmus from blue to red

3. react with .metal (such as Zn, Mg) to produce H_2 gas
- .hydroxide base to produce H_2O and salt
- .carbonate to produce CO_2

Chapter 15 Acids, Bases and Salts

Ch15-2 Bronsted Acids and Bases

An acid is a substance capable of donating a proton, a base is a substance capable of accepting a proton

Concept of a conjugative acid-base pair

$$HCl(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + Cl^-(aq)$$

$$CH_3COOH(aq) + H_2O(l) \rightleftharpoons CH_3COO^-(aq) + H_3O^+(aq)$$

$$NH_3(aq) + H_2O(l) \rightleftharpoons NH_4^+(aq) + OH^-(aq)$$

Chapter 15. Acids and Bases

Chapter 15: Acids and Bases

Acids and Bases Arrhenius Definitions:

acids - compounds that produce an increase in $[H^+]$ when dissolved in water

bases - compounds that produce an increase in $[OH^-]$ when dissolved in water

Lewis Definitions:

acids - electron pair acceptors

bases - electron pair donors

Brønsted-Lowry Definitions:

acids - H^+ donors

Chapter 15: Acids and Bases

Acids and Bases

15 Drill: Identify the B-L acid and base in each of the following.

Circle any amphoteric species

acid acid base base

Note: In both examples, water behaved as an acid or a base. A species that can act as an acid or a base is called amphoteric.

Bronsted-Lowry (B-L) Theory - Cont.

- 1) $HNO_3 + CO_3^{2-} \rightarrow NO_3^- + HCO_3^-$
- 2) $HPO_4^{2-} + H_2O \rightarrow H_2PO_4^-$

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Chapter 15 Acids, Bases and Salts

In the chapter on acids and bases, we saw two more definitions of acids: a compound that donates a proton (a hydrogen ion, H^+) to another compound is called a Brønsted-Lowry acid, and a Lewis acid is any species that can accept a pair of electrons. Explain why the introductory definition is a macroscopic definition, while the Brønsted-Lowry definition and the Lewis definition are microscopic definitions.

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The following acid-base reactions go completely in the forward direction: 1. $HCl(aq) + NH_3(aq) \rightarrow NH_4^+(aq) + Cl^-(aq)$; (HCl is a strong acid) 2. $HSO_4^-(aq) + CN^-(aq) \rightarrow HCN(aq) + SO_4^{2-}(aq)$; (HSO_4^- is a stronger than HCN) Many acid-base reactions reach a state of equilibrium.

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An acid-base reaction occurs in aqueous solution between hydro- chloric acid, a strong acid that completely dissociates to produce H_3O^+ , and sodium hydroxide, a strong base that completely dissociates to produce OH^- . The formula equation for this reaction is written as follows. $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$

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Acid-Base Reactions Reactions of acids and bases a) Reaction of acids and bases with metals. Acid + active metal \rightarrow salt + hydrogen + heat. $2HCl + Mg \rightarrow MgCl_2 + H_2$

$2(↑) \text{Base} + \text{metal} \rightarrow \text{salt} + \text{hydrogen} + \text{heat}$. $2NaOH + Zn \rightarrow Na_2ZnO_2 + H_2(↑)$
 A more reactive metal displaces the less reactive metal from its base. $2Na + Mg(OH)_2 \rightarrow 2NaOH + Mg$

Chapter 15 Acids and Bases - Bridgewater State University

Sec 15.2: The Nature of Acids and Bases: Properties of Acids-sour taste, ability to dissolve many metals, ability to turn blue litmus paper red, and ability to neutralize bases. Properties of Bases-bitter taste, slippery feel, ability to turn red litmus paper blue, and ability to neutralize acids.

Chapter 15 2 Acids Bases

In the chapter on acids and bases, we saw two more definitions of acids: a compound that donates a proton (a hydrogen ion, H^+) to another compound is called a Brønsted-Lowry acid, and a Lewis acid is any species that can accept a pair of electrons. Explain why the introductory definition is a macroscopic definition, while the Brønsted-Lowry definition and the Lewis definition are microscopic definitions.

15.2 Lewis Acids and Bases - General Chemistry 1 & 2

2 15.1 Acids and bases acid derived from Latin acidus (meaning sour or tart) related to Latin acetum (meaning vinegar) characteristic properties associated with acid: 1. sour taste 2. change the color of litmus from blue to red 3. react with .metal (such as Zn, Mg) to produce H_2 gas .hydroxide base to produce H_2O and salt .carbonate to produce CO_2

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Chapter 15. Acids and Bases

Chapter 15 Acids and Bases Multiple Choice Questions 1) _____ is found in carbonated beverages due to the reaction of carbon dioxide with water. A) CH_3COOH ; B) H_2CO_3 ; C) HCOOH ; D) $\text{C}_6\text{H}_5\text{COOH}$; E) $\text{CH}_3\text{CH}_2\text{COOH}$; Answer: B. Diff: 1 Page Ref: 15.2 2) Which of the following species is amphoteric? A) CO_3^{2-} ; B) HF ; C) NH_4^+ ; D) HPO_4^{2-}

Chapter 15: Acids and Bases Acids and Bases

Chapter 15: Acids and Bases Acids and Bases Arrhenius Definitions: acids - compounds that produce an increase in $[\text{H}^+]$ when dissolved in water bases - compounds that produce an increase in $[\text{OH}^-]$ when dissolved in water Lewis Definitions: acids - electron pair acceptors bases - electron pair donors Brønsted-Lowry Definitions: acids - H^+ donors

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15.2 Lewis Acids and Bases - Chemistry

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15 Drill: Identify the B-L acid and base in each of the following. Circle any amphoteric species acid acid base base Note: In both examples, water behaved as an acid or a base. A species that can act as an acid or a base is called amphoteric. Bronsted-Lowry (B-L) Theory - Cont. acid base 1) $\text{HNO}_3 + \text{CO}_3^{2-} \rightleftharpoons \text{NO}_3^- + \text{HCO}_3^-$ 2) $\text{HPO}_4^{2-} + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{PO}_4^-$